

## Assessment

Names and formulas of compounds

Page 73

1. F 2. C 3. I 4. B 5. C 6. D 7. A 8. C 9. D 10. D 11. C  
12. B

## Section 4.3 Chemical Equations

### Comprehension

#### Balancing equations

Page 77

- $\text{H}_2 + \text{F}_2 \rightarrow 2 \text{HF}$
- $2 \text{Sn} + \text{O}_2 \rightarrow 2 \text{SnO}$
- $\text{MgCl}_2 \rightarrow \text{Mg} + \text{Cl}_2$
- $2 \text{KNO}_3 \rightarrow 2 \text{KNO}_2 + \text{O}_2$
- $2 \text{BN} + 3 \text{F}_2 \rightarrow 2 \text{BF}_3 + \text{N}_2$
- $\text{CuI}_2 + \text{Fe} \rightarrow \text{FeI}_2 + \text{Cu}$
- $2 \text{Li} + 2 \text{H}_2\text{O} \rightarrow 2 \text{LiOH} + \text{H}_2$
- $4 \text{NH}_3 + 3 \text{O}_2 \rightarrow 2 \text{N}_2 + 6 \text{H}_2\text{O}$
- $\text{V}_2\text{O}_5 + 5 \text{Ca} \rightarrow 5 \text{CaO} + 2 \text{V}$
- $2 \text{C}_9\text{H}_6\text{O}_4 + 17 \text{O}_2 \rightarrow 18 \text{CO}_2 + 6 \text{H}_2\text{O}$
- $\text{H}_2\text{S} + \text{PbCl}_2 \rightarrow \text{PbS} + 2 \text{HCl}$
- $2 \text{C}_3\text{H}_7\text{OH} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 8 \text{H}_2\text{O}$
- $\text{Zn} + \text{CuSO}_4 \rightarrow \text{Cu} + \text{ZnSO}_4$
- $\text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
- $\text{C}_2\text{H}_5\text{OH} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 3 \text{H}_2\text{O}$
- $2 \text{Al} + 3 \text{H}_2\text{SO}_4 \rightarrow 3 \text{H}_2 + \text{Al}_2(\text{SO}_4)_3$
- $2 \text{FeCl}_3 + 3 \text{Ca}(\text{OH})_2 \rightarrow 2 \text{Fe}(\text{OH})_3 + 3 \text{CaCl}_2$
- $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{PbCrO}_4 + 2 \text{KNO}_3$
- $\text{Cd}(\text{NO}_3)_2 + (\text{NH}_4)_2\text{S} \rightarrow \text{CdS} + 2 \text{NH}_4\text{NO}_3$
- $\text{Ca}(\text{OH})_2 + 2 \text{NH}_4\text{Cl} \rightarrow 2 \text{NH}_3 + \text{CaCl}_2 + 2 \text{H}_2\text{O}$

### Applying Knowledge

#### Word equations

Page 78

- $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
- $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow 3 \text{H}_2\text{O} + 2 \text{Fe}$
- $2 \text{Na} + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + \text{H}_2$
- $\text{Ca}_2\text{C} + \text{O}_2 \rightarrow 2 \text{Ca} + \text{CO}_2$
- $2 \text{KI} + \text{Cl}_2 \rightarrow 2 \text{KCl} + \text{I}_2$
- $4 \text{Cr} + 3 \text{SnCl}_4 \rightarrow 4 \text{CrCl}_3 + 3 \text{Sn}$
- $\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$
- $\text{ZnSO}_4 + \text{SrCl}_2 \rightarrow \text{ZnCl}_2 + \text{SrSO}_4$
- $3 \text{NH}_4\text{Cl} + \text{Pb}(\text{NO}_3)_2 \rightarrow 3 \text{NH}_4\text{NO}_3 + \text{PbCl}_2$
- $2 \text{Fe}(\text{NO}_3)_3 + 3 \text{MgS} \rightarrow \text{Fe}_2\text{S}_3 + 3 \text{Mg}(\text{NO}_3)_2$
- $2 \text{AlCl}_3 + 3 \text{Na}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + 6 \text{NaCl}$
- $2 \text{Na}_3\text{PO}_4 + 3 \text{Ca}(\text{OH})_2 \rightarrow 6 \text{NaOH} + \text{Ca}_3(\text{PO}_4)_2$

## Extension

### Chemical reactions and chemical equations

Page 79

- iron + oxygen  $\rightarrow$  iron(II) oxide  
 $2\text{Fe} + \text{O}_2 \rightarrow 2 \text{FeO}$
- hydrogen chloride + sodium carbonate  $\rightarrow$  carbon dioxide + sodium chloride + water  
 $2 \text{HCl} + \text{Na}_2\text{CO}_3 \rightarrow \text{CO}_2 + 2 \text{NaCl} + \text{H}_2\text{O}$
- aluminum + oxygen  $\rightarrow$  aluminum oxide  
 $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
- water + sodium oxide  $\rightarrow$  sodium hydroxide  
 $\text{H}_2\text{O} + \text{Na}_2\text{O} \rightarrow 2 \text{NaOH}$
- hydrogen + nitrogen trifluoride  $\rightarrow$   
nitrogen + hydrogen fluoride  
 $3 \text{H}_2 + 2 \text{NF}_3 \rightarrow \text{N}_2 + 6 \text{HF}$
- chromium(III) sulphate + potassium carbonate  $\rightarrow$   
chromium(III) carbonate + potassium sulphate  
 $\text{Cr}_2(\text{SO}_4)_3 + 3 \text{K}_2\text{CO}_3 \rightarrow \text{Cr}_2(\text{CO}_3)_3 + 3 \text{K}_2\text{SO}_4$
- potassium chlorate  $\rightarrow$  oxygen + potassium chloride  
 $2 \text{KClO}_3 \rightarrow 3 \text{O}_2 + 2 \text{KCl}$
- zinc + copper(II) sulphate  $\rightarrow$  copper + zinc sulphate  
 $\text{Zn} + \text{CuSO}_4 \rightarrow \text{Cu} + \text{ZnSO}_4$

### Assessment

#### Chemical equations

Page 80

1. B 2. A 3. E 4. D 5. F 6. C 7. G 8. D 9. D 10. D 11. A  
12. C 13. B

## Chapter 5 Compounds are classified in different ways.

### Section 5.1 Acids and Bases

#### Applying Knowledge

#### pH scale and pH indicators

Page 84

- (a) chemical that changes colour depending on the pH of the solution it is placed in  
(b) number scale for measuring how acidic or basic a solution is
- (a)

Substance	pH Value	Acid or Base	Methyl Orange	Bromothymol Blue	Litmus
lemon	2	acid	red	yellow	red
ammonia	11	base	yellow	blue	blue
milk	6	acid	yellow	yellow	red

(b)

Substance	pH Value	Acid or Base	Methyl Red	Phenolphthalein	Indigo Carmine
tomato	4	acid	red	colourless	blue
oven cleaner	13	base	yellow	pink	yellow
egg	8	base	yellow	colourless	blue

3.

Substance	pH Value	Acid or Base	pH Indicator	Colour of pH Indicator
black coffee	5	acid	litmus	red
milk of magnesia	10	base	phenolphthalein	pink
battery acid	0	acid	bromothymol blue	yellow
sea water	8	base	indigo carmine	blue
orange juice	3	acid	methyl orange	red
liquid drain cleaner	14	base	methyl red	yellow

### Comprehension

#### Names of acids

##### Page 86

- ate
- ite
- (a) carbonic acid  
(b) acetic acid  
(c) phosphoric acid  
(d) chlorous acid  
(e) sulphurous acid  
(f) nitric acid  
(g) hydrofluoric acid  
(h) hydrochloric acid
- (a) HI  
(b)  $\text{H}_2\text{SO}_4$   
(c)  $\text{HClO}_4$   
(d)  $\text{HNO}_2$   
(e)  $\text{HClO}_3$   
(f) HBr  
(g)  $\text{H}_3\text{PO}_3$   
(h) HClO

### Applying Knowledge

#### Acids versus bases

##### Page 87

	ACIDS	BASES
definition	compounds containing hydrogen that produce a solution with a pH of less than 7 when they dissolve in water and that produce a salt and water when they react with ionic compounds containing hydroxide ions	chemical compounds containing hydroxide that produce a solution with a pH of more than 7 when they dissolve in water and produce a salt and water when they react with ionic compounds containing positive hydrogen ions
pH	< 7	> 7
what to look for in chemical formula	H	OH
production of ions	$\text{H}^+$	$\text{OH}^-$
electrical conductivity	conductive	conductive
taste	taste sour	taste bitter
touch	burn skin	feel slippery; burn skin
examples	HCl, $\text{H}_2\text{SO}_4$ , lemons, stomach acid	NaOH, KOH, drain cleaner, soap

- (a) acid  
(b) base  
(c) base  
(d) acid  
(e) base  
(f) acid  
(g) acid  
(h) base  
(i) acid  
(j) base  
(k) base  
(l) acid

### Assessment

#### Acids and bases

##### Page 88

- D
- F
- A
- A
- E
- B
- G
- C
- A
- A
- C
- A
- C
- B
- B

## Section 5.2 Salts

### Comprehension

#### Recognizing acids, bases, and salts

##### Page 91

- (a) acid  
(b) acid  
(c) base  
(d) acid  
(e) base  
(f) acid  
(g) acid  
(h) acid  
(i) salt  
(j) base  
(k) base  
(l) salt  
(m) acid  
(n) salt  
(o) salt  
(p) salt  
(q) acid  
(r) acid  
(s) base  
(t) acid  
(u) acid  
(v) salt

- acetic acid,  $\text{CH}_3\text{COOH}$
- sodium chloride,  $\text{NaCl}$
- sulphuric acid,  $\text{H}_2\text{SO}_4$
- sodium hydroxide,  $\text{NaOH}$
- magnesium hydroxide,  $\text{Mg}(\text{OH})_2$
- hydrochloric acid,  $\text{HCl}$

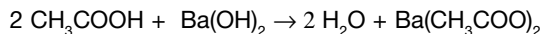
### Applying Knowledge

#### Acid-base neutralization reactions

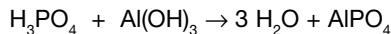
##### Page 92

- (a)  $\text{H}_2\text{SO}_4 + 2 \text{NaOH} \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$   
(b)  $\text{HNO}_3 + \text{KOH} \rightarrow \text{H}_2\text{O} + \text{KNO}_3$   
(c)  $2 \text{HCl} + \text{Ca}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{CaCl}_2$   
(d)  $2 \text{H}_3\text{PO}_4 + 3 \text{Ba}(\text{OH})_2 \rightarrow 6 \text{H}_2\text{O} + \text{Ba}_3(\text{PO}_4)_2$   
(e)  $\text{CH}_3\text{COOH} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCH}_3\text{COO}$   
(f)  $2 \text{HNO}_3 + \text{Sr}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{Sr}(\text{NO}_3)_2$   
(g)  $3 \text{HF} + \text{Fe}(\text{OH})_3 \rightarrow 3 \text{H}_2\text{O} + \text{FeF}_3$   
(h)  $4 \text{HBr} + \text{Sn}(\text{OH})_4 \rightarrow 4 \text{H}_2\text{O} + \text{SnBr}_4$
- (a) sulphuric acid + potassium hydroxide  $\rightarrow$   
water + potassium sulphate  
 $\text{H}_2\text{SO}_4 + 2 \text{KOH} \rightarrow 2 \text{H}_2\text{O} + \text{K}_2\text{SO}_4$

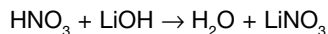
(b) acetic acid + barium hydroxide  $\rightarrow$   
water + barium acetate



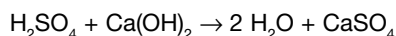
(c) phosphoric acid + aluminum hydroxide  $\rightarrow$   
water + aluminum phosphate



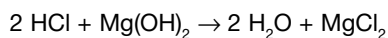
(d) nitric acid + lithium hydroxide  $\rightarrow$   
water + lithium nitrate



(e) sulphuric acid + calcium hydroxide  $\rightarrow$   
water + calcium sulphate



(f) hydrochloric acid + magnesium hydroxide  $\rightarrow$   
water + magnesium chloride



### Applying Knowledge

#### Metal oxides and non-metal oxides

##### Page 93

- oxygen
- metal oxide
- non-metal oxide
- it becomes basic
- it becomes acidic
- a base
- an acid
- (a) metal oxide  
(b) non-metal oxide  
(c) non-metal oxide  
(d) metal oxide  
(e) non-metal oxide  
(f) metal oxide  
(g) non-metal oxide  
(h) metal oxide
- (a) a base  
(b) an acid  
(c) a base  
(d) an acid

### Assessment

#### Salts

##### Page 94

- A 2. C 3. F 4. E 5. D 6. B 7. C 8. B 9. D 10. B 11. B  
12. D 13. B

## Section 5.3 Organic Compounds

### Cloze Activity

#### Organic chemistry

##### Page 98

1. organic compounds; organic chemistry
2. inorganic compounds
3. carbon
4. hydrocarbons
5. methane
6. ethane
7. propane
8. butane
9. alcohol; oxygen
10. solvent
11. ethanol

### Comprehension

#### Recognizing organic and inorganic compounds

##### Page 99

1. inorganic
2. organic
3. inorganic
4. inorganic
5. inorganic
6. inorganic
7. organic
8. organic
9. organic
10. organic
11. inorganic
12. inorganic
13. organic
14. inorganic
15. organic
16. organic
17. inorganic
18. organic
19. inorganic
20. inorganic
21. inorganic
22. inorganic
23. organic
24. inorganic
25. organic
26. organic
27. organic
28. organic

29. organic

30. organic

### Applying Knowledge

#### Organic compounds versus inorganic compounds

##### Page 100

1. organic
2. organic
3. inorganic
4. organic
5. organic
6. organic
7. inorganic
8. organic

### Assessment

#### Organic compounds

##### Page 101

1. B 2. A 3. D 4. C 5. A 6. D 7. D 8. A

## Chapter 6 Chemical reactions occur in predictable ways.

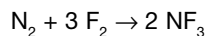
### Section 6.1 Types of Chemical Reactions

#### Comprehension

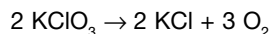
#### Classifying chemical reactions

##### Page 105

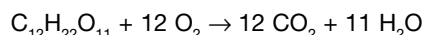
1. S



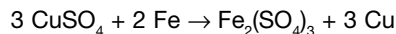
2. D



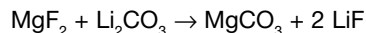
3. C



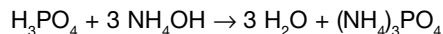
4. SR



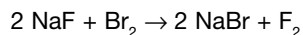
5. DR



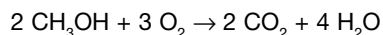
6. N



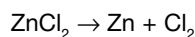
7. SR



8. C



9. D



10. DR

